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Elements, Metric
System \u0026

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Introduction to
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Chemical

Equation

Calculations

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Stoichiometry

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Section 2 covers

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including mole
to mole, mole to
mass, mass to
mole, and mass
to mass
examples.

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Section 2:

Stoichiometric Calculations ...

Stoichiometry.

SECTION 2.

PROBLEMS Write the answer on the line to the left. Show all your work in the space provided.

1. The following equation

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book with a cup
of tea in the
afternoon,
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relationships of
elements in
compounds
Reaction

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Involves mass
relationships
between
reactants and
products in a
chemical
reaction I. Page
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office, home,
and other
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you go. So, you
won't have

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carry. This is
why your choice
to make ...

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the tape sooner.

But, sometimes
it is thus in
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city. So, to
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the answer on

the line to the

left. Show all

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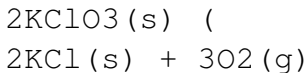
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your work in the
space provided.

1. The following
equation

represents a
laboratory
preparation for
oxygen gas:



How many moles
of O_2 form if
3.0 mol of KClO_3
are totally

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Stoichiometry

Section 2

Worksheet

consumed? 2.

Given the
following

equation: $\text{H}_2(\text{g})$

+ $\text{F}_2(\text{g}) \rightarrow 2\text{HF}(\text{g})$

CHAPTER 9 REVIEW

Chapter menu

Resources

Chapter 9

Section 1

Introduction to

Stoichiometry

Objective •

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Stoichiometry

Define stoichiometry. •

Describe the importance of the mole ratio in

stoichiometric calculations. •

Write a mole ratio relating two substances in a chemical equation.

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calculations

involving the

mass

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relationships

between

reactants and

products in a

chemical

reaction . Mole

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SECTION 3

PROBLEMS Write

the answer on

the line to the

left. Show all

your work in the

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Stoichiometry.

1. 88% The actual yield of a reaction is 22

g and the theoretical yield is 25 g.

Calculate the percentage

yield. 2. 6.0

mol of N_2 are mixed with 12.0

mol of H_2

according to the

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Chapter 9

Stoichiometry

equation: $\text{N}_2(\text{g})$

$3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$

N_2 ; 2.0

mol a. Which

chemical is ...

*mc06se cFMsr i-
vi*

SECTION 1

Introduction to
Stoichiometry

SECTION 2 Ideal
Stoichiometric

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Chapter 9

Stoichiometry

SECTION 3

Section 2

Limiting

Worksheet

Reactants and

Percentage Yield

Why It Matters

Video

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Stoichiometry

BIG IDEA

Reaction

stoichiometry

uses molar

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Chapter 9

Stoichiometry to
determine the
amounts of
unknown

reactants or
products from
the amounts of
known reactants
or products.

CHAPTER 9 DO NOT
EDIT--Changes
must be ...

CorrectionKey=NL-
Page 47/55

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Chapter 9

Stoichiometry

EDIT--Changes

must be made ...

Chapter 9.

Solutions.

Search for: 9.3

Solution

Stoichiometry.

Learning

Objectives. By

the end of this

section, you

will be able to:

Perform

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Chapter 9

Stoichiometric

calculations

involving

solution

molarity; As we

have seen in

lab, many

reactions such

as single or

double

displacement

reactions are

carried out in

aqueous medium

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Chapter 9

(i.e. in water) .

Because these reactions occur in aqueous solution, we can use ...

9.3 Solution

Stoichiometry /

Introductory

Chemistry

Chapter 9

STOICHIOMETRY.

What is Sto-Key-

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Chapter 9

Om-It-Tree?

Stoichiometry is
derived from the
Greek words

stoicheion

“element” and

metron

“measure”.

Stoichiometry is
defined as the
branch of
chemistry that
deals with the
mass

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Stoichiometry
relationships
between elements
in a compound
and the mass
relationships
between
reactants and
products in a
chemical
reaction. There
are 2 main types
of stoichiometry
...

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Stoichiometry -

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11.2. As you

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section,

summarize each

step on a tab

and include an

example of the

step. 368

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Worksheet
Objectives

Describe the types of relationships indicated by a balanced chemical equation. State the mole ratios from a balanced chemical

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Stoichiometry Review

Vocabulary

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reactant: the
starting

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